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## **Stoichiometry Problem Sheet 1 Answers**

Chemistry: Stoichiometry – Problem

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Sheet 1 Directions: Solve each of the following problems. Show your work, including proper units, to earn full credit.

1. Silver and nitric acid react according to the following balanced equation:  $3 \text{Ag}(s) + 4 \text{HNO}_3(aq) \rightarrow 3 \text{AgNO}_3(aq) + 2 \text{H}_2\text{O}(l) + \text{NO}(g)$  A.

## **Stoichiometry: Problem Sheet 1**

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Stoichiometry WorkSheet #1: Worked Solutions Answer the following questions on your own paper. Show all work. Circle the final answer, giving units and the correct number of significant figures. 1. Based on the following equation, how many moles of each product are produced when 5.9 moles of  $\text{Zn(OH)}_2$  are reacted with  $\text{H}_3\text{PO}_4$ ? (You need

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## **Stoichiometry WorkSheet #1: Worked Solutions**

Stoichiometry: Problem Sheet 1 -  
teachnlearnchem.com (ANSWER 386.3g  
of  $\text{LiNO}_3$ ) 4) Using the following  
equation:  $\text{Fe}_2\text{O}_3 + 3 \text{H}_2 \rightarrow 2 \text{Fe} + 3 \text{H}_2\text{O}$ . Calculate how many grams of iron  
can be made from 16.5 grams of  $\text{Fe}_2\text{O}_3$

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... Unit 08 - Stoichiometry - Worksheet 1  
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of stoichiometry with tons of interesting examples and exercises involving with step by step solutions with several colorful illustrations and diagrams.

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Stoichiometry Practice #1 KEY - chemistrygods.net. chemistrygods.net.

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Stoichiometry Practice Worksheet. 13)  
Using the equation from problem #12,  
determine the mass of aluminum  
acetate that can be made if I do this  
reaction with 125 grams of acetic acid  
and 275 grams of aluminum hydroxide.

## **Stoichiometry Practice Problems**

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## Level 1 Answers

(ANSWER 386.3g of  $\text{LiNO}_3$ ) 4) Using the following equation:  $\text{Fe}_2\text{O}_3 + 3 \text{H}_2 \rightarrow 2 \text{Fe} + 3 \text{H}_2\text{O}$ . Calculate how many grams of iron can be made from 16.5 grams of  $\text{Fe}_2\text{O}_3$  by the following equation. Worksheet for Basic Stoichiometry. Part 1: Mole  $\leftrightarrow$  Mass Conversions. Convert the following

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number of moles of chemical into its corresponding mass in grams.

## **Worksheet for Basic Stoichiometry**

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mol  $\text{KClO}_3$  3 mol  $\text{KClO}_3$  mol  $\text{O}_2$  = mol  $\text{O}_2$   
2 3.50 mol  $\text{KCl}$  = mol  $\text{KClO}_3$  = 0.275 mol  
 $\text{Fe} = \text{mol Fe}_2\text{O}_3 = = 2 \text{KClO}_3 \rightarrow 2 \text{KCl} +$   
3  $\text{O}_2$  10 ...

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Stoichiometry Practice Worksheet Solve  
the following stoichiometry grams-grams  
problems: 1) Using the following  
equation:  $2 \text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} +$

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Na<sub>2</sub>SO<sub>4</sub> 4 How many grams of sodium sulfate will be formed if you start with 200.0

## **Stoichiometry Worksheet 1 Answer Key**

Learning Target Perform mole-to-mole conversions Stoichiometry worksheet 1 mole to mole calculations answers.

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Directions : You must solve each of the following problems using dimensional analysis. 1. For this reaction:  $\underline{\quad}$ Al +  $\underline{\quad}$ O<sub>2</sub> →  $\underline{\quad}$ Al<sub>2</sub>O<sub>3</sub> a. How many moles of aluminum oxide will be formed from 17 moles of aluminum reacting? b.

## **Stoichiometry Worksheet 1 Mole To Mole Calculations Answers**

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## **Chemistry Stoichiometry Problem Sheet 1 Answers [EPUB]**

Stoichiometry Worksheet #1 Answers 1.  
Given the following equation:  $2 \text{C}_4\text{H}_{10} + 13 \text{O}_2 \rightarrow 8 \text{CO}_2 + 10 \text{H}_2\text{O}$ , show  
what the following molar ratios should

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be. a. C 4H 10 / O 2 b. O 2 / CO 2 c. O 2 /  
H 2O d. C 4H 10 / CO 2 e. C 4H 10 / H 2O  
2. Given the following equation:  $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$   
a. How many moles of O 2 can be produced by ...

## **Stoichiometry Worksheet #1 Answers - PSD401**

Stoichiometry Volume Problems



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## Worksheet Answers

Stoichiometry Volume Problems

Worksheet Answers Stoichiometry:

Mixed Problems (KEY) 1)  $N_2 + 3H_2 \rightarrow$

$2NH_3$  What volume of  $NH_3$  at STP is produced if 25.0 of  $N_2$  is reacted with an excess of  $H_2$ ? 3 3 3 2 3 2 2 2 40.0L  $NH_3$

1mol  $NH_3$  22.4L  $NH_3$  1mol  $N_2$  2mol  $NH_3$

28.0g  $N_2$  25.0g  $N_2$  1mol  $N_2$   $\times \times \times = 2) \dots$

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## **Stoichiometry Volume Problems Worksheet Answers**

Chemistry: Stoichiometry – Problem  
Sheet 2 KEY 9) 2 24 2 2 23 2 2 2 2 4.63  
 $\times 10$  molecules | 1 mol |  $6.02 \times 10$   
molecules | 1 mol Cl | 1 mol 71 g Cl Cl  $\times$   
546 g Cl 10) 292 g Ag 1 mol Ag 108 g Ag  
1 mol Cu 1 mol Ag 63.5 g Cu

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